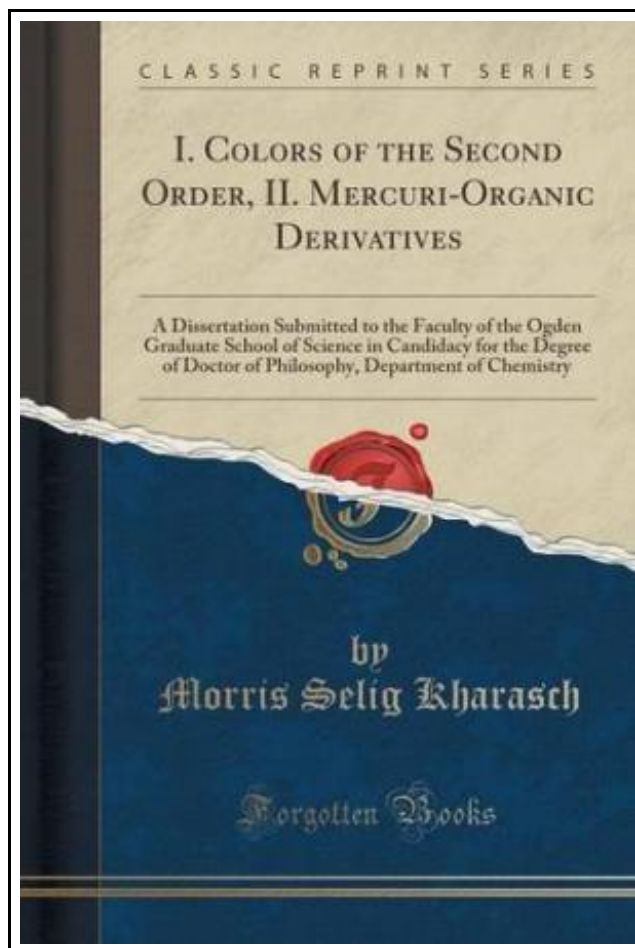


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Forgotten Books, United States, 2015. Paperback. Book Condition: New. 229 x 152 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.Excerpt from **I. Colors of the Second Order, II. Mercuri-Organic Derivatives: A Dissertation Submitted to the Faculty of the Ogden Graduate School of Science in Candidacy for the Degree of Doctor of Philosophy, Department of Chemistry** It is a well-known rule that the color of a yellow dye changes with increase of molecular weight to orange, red, violet, blue, and finally to green. Each one of these colors is called lower than the preceding one, and higher than the following one. Generally, we can lower the color by the introduction of alkyl or aryl radicles into the auxochromic or chromophoric groups of the dye. The spectroscopic explanation of this behavior of dyes is that a yellow dye has an absorption band in the violet, and any increase in the molecular weight decreases the number of vibrations corresponding to this absorption band causing the band to shift toward the red end of the spectrum. When the absorption band is in the blue the compound appears orange, and as the absorption band goes through green, yellow, orange to red, the color of the respective compounds successively changes to red, violet, blue, and green. A compound has the color complementary to that of its maximum absorption. On the other hand, if in the molecule of a yellow dye we make such changes as generally tend to produce a higher color, then the absorption bands shift into the ultra-violet part of the spectrum, the compound becoming colorless. Absorption colors have very much in common with interference colors. The colors of Newton, for instance, are also produced by an absorption band going through the spectrum in the same direction in which the...

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